



THE UNIVERSITY *of* EDINBURGH

## Edinburgh Research Explorer

# Iron-catalysed, hydride-mediated reductive cross-coupling of vinyl halides and Grignard reagents

### Citation for published version:

Le Bailly, BAF, Greenhalgh, MD & Thomas, SP 2012, 'Iron-catalysed, hydride-mediated reductive cross-coupling of vinyl halides and Grignard reagents', *Chemical Communications*, vol. 48, no. 10, pp. 1580-1582. <https://doi.org/10.1039/c1cc14622j>

### Digital Object Identifier (DOI):

[10.1039/c1cc14622j](https://doi.org/10.1039/c1cc14622j)

### Link:

[Link to publication record in Edinburgh Research Explorer](#)

### Document Version:

Peer reviewed version

### Published In:

Chemical Communications

### Publisher Rights Statement:

Copyright © 2011 by the Royal Society of Chemistry. All rights reserved.

### General rights

Copyright for the publications made accessible via the Edinburgh Research Explorer is retained by the author(s) and / or other copyright owners and it is a condition of accessing these publications that users recognise and abide by the legal requirements associated with these rights.

### Take down policy

The University of Edinburgh has made every reasonable effort to ensure that Edinburgh Research Explorer content complies with UK legislation. If you believe that the public display of this file breaches copyright please contact [openaccess@ed.ac.uk](mailto:openaccess@ed.ac.uk) providing details, and we will remove access to the work immediately and investigate your claim.



Post-print of a peer-reviewed article published by the Royal Society of Chemistry.  
Published article available at: <http://dx.doi.org/10.1039/C1CC14622J>

Cite as:

Le Bailly, B. A. F., Greenhalgh, M. D., & Thomas, S. P. (2012). Iron-catalysed, hydride-mediated reductive cross-coupling of vinyl halides and Grignard reagents. *Chemical Communications*, 48(10), 1580-1582.

Manuscript received: 28/07/2011; Accepted: 14/10/2011; Article published: 31/10/2011

## Iron-catalysed, hydride-mediated reductive cross-coupling of vinyl halides and Grignard reagents<sup>\*,§</sup>

Bryden A. F. Le Bailly, Mark D. Greenhalgh and Stephen P. Thomas<sup>†,\*</sup>

School of Chemistry, University of Bristol, Cantock's Close, Bristol, England BS8 1TS, UK.

[\*]Corresponding author; fax: +44(0)117 925 1295; e-mail: [stephen.thomas@ed.ac.uk](mailto:stephen.thomas@ed.ac.uk)

[\*\*]We thank Prof. V. K. Aggarwal for his support and generosity. SPT thanks N. C. O. Tomkinson, S. Warren and D. J. Fox for continued support.

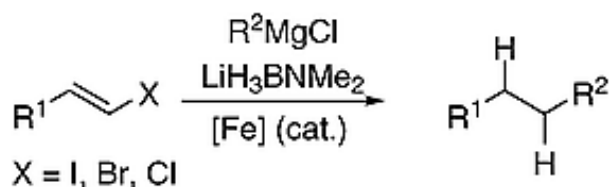
[§]This article is part of a ChemComm 'Emerging Investigators' theme issue showcasing high quality research in organic chemistry. Please go to RSC website at <http://pubs.rsc.org/en/journals/journalissues/cc#!issueid=cc048010&type=current&issnprint=1359-7345> to access the other papers in this issue.

[†]Current affiliation: EaStCHEM, School of Chemistry, Joseph Black Building, University of Edinburgh, West Mains Road, Edinburgh, EH9 3JJ, UK.

### Supporting information:

Electronic Supplementary Information (ESI) available – full experimental procedures and analytical data are available. See <http://dx.doi.org/10.1039/C1CC14622J>

### Graphical abstract:



### Synopsis:

An Iron-catalysed, hydride-mediated reductive cross-coupling has been developed and used to prepare a range of alkanes from vinyl halides and Grignard reagents.

## Abstract

An iron-catalysed, hydride-mediated reductive cross-coupling reaction has been developed for the preparation of alkanes. Using a bench-stable iron(II) pre-catalyst, reductive cross-coupling of vinyl iodides, bromides and chlorides with aryl- and alkyl Grignard reagents successfully gave the products of formal  $sp^3$ - $sp^3$  cross-coupling reactions.

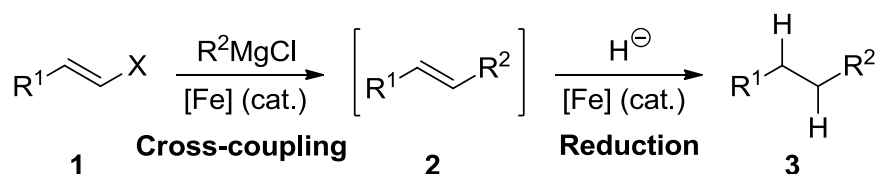
## Introduction

Metal-catalysed cross-coupling reactions have become ubiquitous in synthetic chemistry with palladium- and nickel-catalysed reactions used routinely in both academia and industry.<sup>1</sup> However, preceding the seminal discoveries of Heck<sup>2</sup> and Kumada,<sup>3</sup> Kharasch<sup>4</sup> and, later, Kochi reported the iron-catalysed cross-coupling of vinyl bromides and Grignard reagents.<sup>5</sup> Iron-catalysed cross-coupling reactions have since been developed into a powerful method for the formation of carbon-carbon<sup>6</sup> and carbon-heteroatom bonds.<sup>7</sup>

The last decade has seen a significant increase in the number of transformations achieved using iron catalysts.<sup>8</sup> The iron-catalysed reduction of carbonyl groups by hydrosilylation, transfer-hydrogenation and hydrogenation are now well developed, with turnover frequencies equal to those of the best ruthenium-based catalysts and good enantioselectivities have been achieved for the reduction of ketones and imines.<sup>9</sup> However, the reduction of olefins is only now being realised without the need for forcing reaction conditions. Chirik has developed a series of iron(0) complexes, based on the polymerisation catalysts of Gibson and Brookhart,<sup>10</sup> which catalyse the hydrogenation of functionalised and unfunctionalised alkenes at low hydrogen pressure and with excellent activity.<sup>11</sup> de Vries has used iron nanoparticles as the catalyst for the hydrogenation of alkenes and alkynes.<sup>12</sup> However, both of these powerful methods use air- and moisture-sensitive iron species which require synthesis prior to the hydrogenation reaction. Iron porphyrin complexes and sodium borohydride,<sup>13</sup> and iron chloride in conjunction with lithium aluminium hydride<sup>14</sup> have also been used to reduce alkenes,  $\alpha,\beta$ -unsaturated ketones and esters.

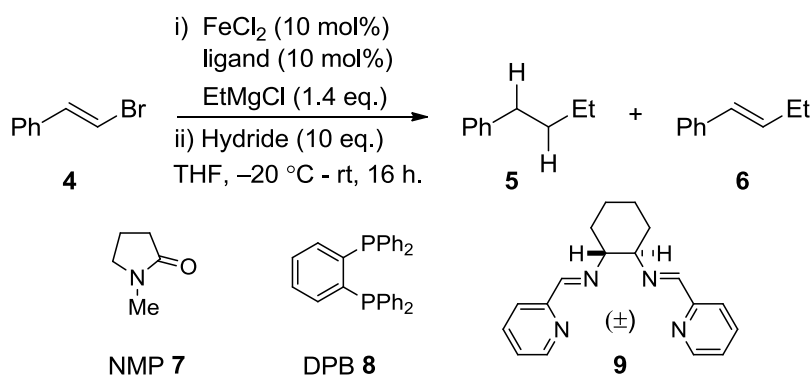
We recently reported the iron-catalysed, hydride-mediated reduction of alkenes using substoichiometric iron triflate and sodium triethylborohydride.<sup>15</sup> Examination of the postulated catalytic cycles of the iron-catalysed hydrogenation<sup>11</sup> and cross-coupling<sup>16</sup> reactions reveal common low-valent iron species. We therefore sought to combine both of these reactions into one transformation, catalysed by a single, bench stable, iron pre-catalyst. Reductive cross-coupling is proposed to proceed by iron-catalysed cross-coupling of the vinyl halide and Grignard reagent to give an alkene **2**. This alkene is then reduced *in situ* by an iron hydride to the alkane **3** (Figure 1). The iron

hydride is generated by the reaction of the iron salt with a stoichiometric hydride reagent. This transformation gives the product of an  $sp^3$ - $sp^3$  cross-coupling without the problems associated with these reactions.<sup>1</sup> Furthermore, as the Grignard reagent needed for cross-coupling generates a low-valent iron species *in situ*, we will be able to use bench-stable iron(II) pre-catalysts.



**Figure 1.** Iron-catalysed, hydride-mediated reductive cross-coupling.

In order to develop the reductive cross-coupling we chose commercially available  $\beta$ -bromostyrene **4** and ethyl magnesium chloride as model coupling partners and iron(II) chloride as the iron source (Table 1). Sodium triethylborohydride was initially used as the hydride source and *N*-methyl pyrrolidinone (NMP) **7** as a stabilising ligand due to the success we had with these reagents for the isolated reduction of alkenes.<sup>15</sup> Using 10 mol% iron(II) chloride, 110 mol% NMP and 10 equivalents of NaHBET<sub>3</sub> gave the reduced cross-coupled alkane **5** in only 9% conversion and the cross-coupled alkene **6** in 53% (entry 1). This indicated that although cross-coupling was successful, the alkene was not being reduced under these conditions. We assumed that either the iron catalyst was not capable of mediating the hydrogenation or was degrading before hydrogenation could occur. To extend catalyst lifetime we screened multidentate ligands and found that the tetradentate dipyridyldiimino ligand **9** gave 56% conversion to the alkane **5** and 44% alkene **6** (entries 2 and 3). This showed that complete cross-coupling was achieved and the hydrogenation step now needed improvement. Thus, we screened a number of hydride sources (entries 4-12). Variation of the cation of the hydride source showed that KHB<sup>s</sup>Bu<sub>3</sub> (K-Selectride) (entry 6) gave a similar conversion as NaHBET<sub>3</sub>, but the use of reagents with a lithium counter ion gave the best conversions (entries 7-9). LiHBET<sub>3</sub> gave a 20% increase in alkane **5** (entry 7), compared to NaHBET<sub>3</sub>, and lithium *N,N*-dimethylaminoborohydride<sup>17</sup> gave quantitative conversion to the alkane **5** (entry 8). It should be noted that even in the absence of ligand the reductive cross-coupling proceeded, but only 61% of the alkane **5** and 24% alkene **6** were observed (Table 1, entry 9). Without iron(II) chloride no cross-coupling or reduction products were observed. To test if the reductive cross-coupling was catalysed by an iron species or a trace metal impurity,<sup>18</sup> the reaction was carried out with added Cu, Ni and Co salts and with these metals in the absence of iron. No reductive cross-coupling was observed using Cu, Ni or Co salts alone and the yield of the reaction was not increased when these salts were added to the standard reaction conditions.<sup>19</sup>



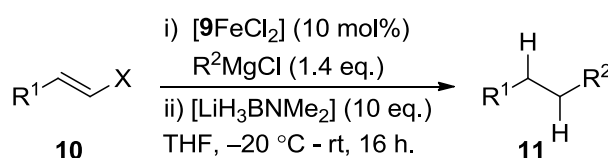
| Entry | Hydride                               | Ligand                    | Conversion (%) <sup>b</sup> |                 |
|-------|---------------------------------------|---------------------------|-----------------------------|-----------------|
|       |                                       |                           | Alkane <b>5</b>             | Alkene <b>6</b> |
| 1     | NaHBET <sub>3</sub>                   | NMP <b>7</b> <sup>c</sup> | 9                           | 53              |
| 2     | NaHBET <sub>3</sub>                   | DPB <b>8</b>              | 2                           | 91              |
| 3     | NaHBET <sub>3</sub>                   | <b>9</b>                  | 56                          | 44              |
| 4     | NaHB(OAc) <sub>3</sub>                | <b>9</b>                  | -                           | 78              |
| 5     | LiHAl(O <sup>t</sup> Bu) <sub>3</sub> | <b>9</b>                  | 25                          | 73              |
| 6     | KHB <sup>s</sup> Bu <sub>3</sub>      | <b>9</b>                  | 55                          | 44              |
| 7     | LiHBET <sub>3</sub>                   | <b>9</b>                  | 73                          | 27              |
| 8     | LiH <sub>3</sub> BNMe <sub>2</sub>    | <b>9</b>                  | 99                          | -               |
| 9     | LiH <sub>3</sub> BNMe <sub>2</sub>    | -                         | 61                          | 24              |
| 10    | NaBH <sub>4</sub>                     | <b>9</b>                  | -                           | 93              |
| 11    | LiAlH <sub>4</sub>                    | <b>9</b>                  | -                           | 94              |
| 12    | HSiEt <sub>3</sub>                    | <b>9</b>                  | -                           | 87              |

<sup>a</sup> Conditions: 1 mmol **4**, 10 mol% FeCl<sub>2</sub>, 10 mol% ligand, THF (0.1 M), -20 °C, i) 1.4 eq. EtMgCl (2M in THF), 15 min, ii) 10 eq. hydride, warm to rt, 16 h. <sup>b</sup> Conversion measured by GC-MS of the crude reaction mixture by comparison to authentic samples. <sup>c</sup> 110 mol% NMP.

**Table 1.** Optimisation of the iron-catalysed, hydride-mediated reductive cross-coupling<sup>a</sup>

Having developed optimised reaction conditions we investigated the scope and limitations of the reductive cross-coupling (Table 2). We began by using β-bromostyrene and a range of alkyl Grignard reagents (entries 1-6). Ethyl- and iso-propylmagnesium chloride both gave excellent conversions to the alkanes (entries 2 and 3). *tert*-Butyl- and phenylmagnesium chloride gave a moderate conversion to the alkane, but no alkene was observed (entries 4-5). Using Grignard reagents without a β-hydrogen was expected to give poor conversion as the formation of the key low-valent iron species is now not possible. However, methylmagnesium chloride gave moderate conversion to the alkane **11a** (entry 1). Presumably an anionic ferrate complex<sup>17</sup> is the catalytically active iron species when using this Grignard reagent. Using allylmagnesium chloride introduced a second alkene to the reaction and this was also reduced under the reaction conditions to give the fully saturated alkane **11e**, albeit in low yield (entry 6).

We next turned our attention to the effect of varying the halide substituent of the vinyl halide (entries 7-12). Using the elegant methodology of Charette<sup>20</sup> we prepared a series of aryl-vinyl halides bearing differing aryl and vinyl functionality.  $\beta$ -Iodo- and  $\beta$ -bromo-4-methylstyrene gave equal conversions to the alkane **11f**, but  $\beta$ -chloro-4-methylstyrene gave 13% of the alkane **11f** and 46% of the alkene (entries 7-9). Electron-deficient aryl-vinyl halides showed no such dependence on the vinyl halide substituent. Both  $\beta$ -chloro- and  $\beta$ -bromo-4-chlorostyrene gave similar conversion to the alkane **5** (entries 10-11). However, proto-dehalogenation of the aromatic ring was observed in both cases, possibly suggesting a low-valent iron catalyst.<sup>21</sup> Finally, we tested the electron-rich  $\beta$ -bromo-3,5-dimethoxystyrene and this gave comparable conversion to that observed when using the electron-neutral aryl-vinyl halides (entry 12).



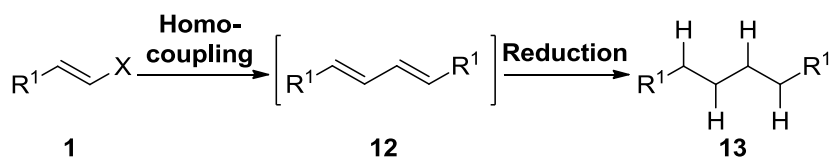
| Entry | R <sup>1</sup> | X  | R <sup>2</sup> | Alkane                | Conversion (%) <sup>b</sup> |
|-------|----------------|----|----------------|-----------------------|-----------------------------|
| 1     | Ph             | Br | Me             | 53                    | <b>11a</b>                  |
| 2     | Ph             | Br | Et             | 99 (84 <sup>c</sup> ) | <b>5</b>                    |
| 3     | Ph             | Br | Pr             | 87                    | <b>11b</b>                  |
| 4     | Ph             | Br | Bu             | 37 <sup>d</sup>       | <b>11c</b>                  |
| 5     | Ph             | Br | Ph             | 41 <sup>e</sup>       | <b>11d</b>                  |
| 6     | Ph             | Br | Allyl          | 35                    | <b>11e<sup>f</sup></b>      |
| 7     | 4-Me-Ph        | I  | Et             | 92 (87 <sup>c</sup> ) | <b>11f</b>                  |
| 8     | 4-Me-Ph        | Br | Et             | 91                    | <b>11f</b>                  |
| 9     | 4-Me-Ph        | Cl | Et             | 13 <sup>g</sup>       | <b>11f</b>                  |
| 10    | 4-Cl-Ph        | Cl | Et             | 60                    | <b>5</b>                    |
| 11    | 4-Cl-Ph        | Br | Et             | 66                    | <b>5</b>                    |
| 12    | 3,5-MeO-Ph     | Br | Et             | 81                    | <b>11g</b>                  |

Conditions: 1 mmol **10**, 10 mol% [9FeCl<sub>2</sub>], THF (0.1 M), -20 °C, i) 1.4 eq. R<sup>2</sup>MgCl (2M in THF), 15 min, ii) 10 eq. [LiH<sub>3</sub>BNMe<sub>2</sub>] (1.7 M in THF), warm to rt, 16 h. <sup>b</sup> Conversion measured by GC-MS of the crude reaction mixture by comparison to authentic samples. <sup>c</sup> Isolated yield. <sup>d</sup> 61% 1,4-diphenylbutane. <sup>e</sup> 40% biphenyl. <sup>f</sup> Complete reduction of R<sup>2</sup> observed in the reaction, product **11e** R<sup>2</sup> = *n*-propyl. <sup>g</sup> 46% alkene.

**Table 2.** Scope and limitation of the iron-catalysed, hydride-mediated reductive cross-coupling<sup>a</sup>

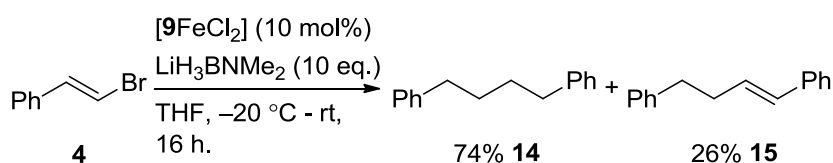
At this point we were unsure why the low alkane conversions were not accompanied by significant amounts of the intermediary alkenes. In the cases with low conversion to the alkane **11**, careful examination of the crude reaction mixtures revealed that the vinyl halide **10** was undergoing reductive

homo-coupling to give the 1,4-diarylbutane **13**, presumably *via* the 1,4-diarylbutadiene **12** (Figure 2).



**Figure 2.** Proposed homocoupling of vinyl halides.

In order to address this we examined the time between Grignard and hydride addition, the time over which the Grignard reagent was added, and the initial temperature of the reaction. Under our standard reaction conditions using <sup>i</sup>PrMgCl, variation of all these parameters led to very similar results with no increase in the conversion to the target alkane. We therefore suspected the hydride reagent was causing the low conversions. Carrying out the reaction in the absence of Grignard reagent confirmed this. Reaction of β-bromostyrene **4** with 10 mol% [9FeCl<sub>2</sub>] and 10 equivalents of lithium *N,N*-dimethylaminoborohydride gave 74% conversion to the fully reduced homo-coupled product, 1,4-diphenylbutane **14**, and 26% of the mono-reduced homo-coupled alkene **15** (Scheme 1). Presumably when cross-coupling is slow, for example with <sup>t</sup>BuMgCl, the rate of hydride-mediated homo-coupling becomes competitive with the rate of Grignard-mediated cross-coupling. Thus, on addition of the hydride reagent any remaining vinyl halide **10** can either undergo reductive homo-coupling to give the diarylbutane **13** or reductive cross-coupling to alkane **11**. Further studies are ongoing to overcome and exploit this problem.



**Scheme 1.** Iron-catalysed, hydride-mediated homo-coupling of β-bromostyrene

In summary, we have developed an iron-catalysed, hydride-mediated reductive cross-coupling reaction for the preparation of alkanes from vinyl halides and Grignard reagents using a simple, bench-stable iron(II) pre-catalyst and commercially available reagents. This transformation gives the product of a formal sp<sup>3</sup>-sp<sup>3</sup> cross-coupling reaction by the *in situ* reduction of an intermediate alkene to the alkane. The scope and limitations of this reaction have been tested and moderate to excellent conversions have been observed in the reductive cross-coupling of vinyl bromides, iodides and chlorides with aryl and alkyl Grignard reagents.

## Notes and references

- [1] F. Diederich, P. J Stang, Eds., *Metal-catalyzed Cross-coupling Reactions*, Wiley-VCH, Weinheim, Germany, 1998.
- [2] R. F. Heck and J. P. Nolley, Jr., *J. Org. Chem.*, 1972, **37**, 2320.
- [3] K. Tamao, K. Sumitani and M. Kumada, *J. Am. Chem. Soc.*, 1972, **94**, 4374.
- [4] M. S. Kharasch and E. K. Fields, *J. Am. Chem. Soc.*, 1941, **63**, 2316.
- [5] M. Tamura and J. K. Kochi, *J. Am. Chem. Soc.*, 1971, **93**, 1487; R. S. Smith and J. K. Kochi, *J. Org. Chem.*, 1976, **41**, 502.
- [6] For reviews see: E. Nakamura and N. Yoshikai, *J. Org. Chem.*, 2010, **75**, 6061; W. M. Czaplik, M. Mayer, J. Cvengros, and A. Jacobi von Wangelin, *ChemSusChem*, 2009, **2**, 396; A. Leitner, In *Iron Catalysis in Organic Chemistry: Reactions and Applications*, B. Pleitker, Ed., Wiley-VCH, Weinheim, Germany, 2008, pp 147-176; A. Rudolph and M. Lautens, *Angew. Chem. Int. Ed.*, 2009, **48**, 2656; B. M. Sherry and A. Fürstner, *Acc. Chem. Res.*, 2008, **41**, 1500; For selected recent examples see: A. K. Steib, T. Thaler, K. Komeyama, P. Mayer and P. Knochel, *Angew. Chem. Int. Ed.*, 2011, **50**, 3303; W. M. Czaplik, M. Mayer and A. Jacobi von Wangelin, *ChemCatChem*, 2011, **3**, 135; M. Mayer, W. M. Czaplik and A. Jacobi von Wangelin, *Adv. Synth. Catal.*, 2010, **352**, 2147; R. B. Bedford, M. Huwe and M. C. Wilkinson, *Chem. Commun.* 2009, 600; G. Cahiez, V. Habiak, C. Duplais and A. Moyeux, *Angew. Chem. Int. Ed.*, 2007, **46**, 4364.
- [7] A. Correa, O. García Mancheño and C. Bolm, *Chem. Soc. Rev.*, 2008, **37**, 1108; M. S. Chen and M. C. White, *Science*, 2010, **327**, 566; K. Suzuki, P. D. Oldenburg and L. Que Jr., *Angew. Chem. Int. Ed.*, 2008, **47**.
- [8] *Iron Catalysis in Organic Chemistry: Reactions and Applications*; B. Pleitker, Ed., Wiley-VCH, Weinheim, Germany 2008; O. G. Mancheño, *Angew. Chem. Int. Ed.*, 2011, **50**, 2216; S. Enthaler, K. Junge and M. Beller, *Angew. Chem. Int. Ed.*, 2008, **47**, 3317.
- [9] For reviews see: K. Junge, K. Schröder and M. Beller, *Chem. Commun.*, 2011, **47**, 4849; P. J. Chirik, In *Catalysis without Precious Metals*, R. M. Bullock, Ed., Wiley-VCH: Weinheim, Germany, 2010, pp 83-110; For selected recent examples see: R. Langer, G. Leitus, Y. Ben-David and D. Milstein, *Angew. Chem. Int. Ed.*, 2011, **50**, 2120; P. O. Lagaditis, A. J. Lough and R. H. Morris, *J. Am. Chem. Soc.*, 2011, **133**, 9662; R. H. Morris, *Chem. Soc. Rev.*, 2009, **38**, 2282; S. Enthaler, A. M. Tondreau, J. M. Darmon, B. M. Wile, S. K. Floyd, E. Lobkovsky and P. J. Chirik, *Organometallics*, 2009, **28**, 3928; K. Junge and M. Beller, *Angew. Chem. Int. Ed.*, 2008, **47**, 3317; C. P. Casey and H. Guan, *J. Am. Chem. Soc.* 2007, **129**, 5816.



- [10] G. J. P. Britovsek, V. C. Gibson, B. S. Kimberley, P. J. Maddox, S. J. McTavish, G. A. Solan, A. J. P. White and D. J. Williams, *Chem. Commun.*, 1998, 849; B. L. Small, M. Brookhart and A. M. A. Bennett, *J. Am. Chem. Soc.*, 1998, **120**, 4049.
- [11] S. C. Bart, E. Lobkovsky and P. J. Chirik, *J. Am. Chem. Soc.*, 2004, **126**, 13794; S. C. Bart, E. J. Hawrelak, E. Lobkovsky and P. J. Chirik, *Organometallics*, 2005, **24**, 5518; R. J. Trovitch, E. Lobkovsky and P. J. Chirik, *Inorg. Chem.*, 2006, **45**, 7252; S. C. Bart, K. Chlopek, E. Bill, M. W. Bouwkamp, E. Lobkovsky, F. Neese, K. Wiegardt and P. J. Chirik, *J. Am. Chem. Soc.*, 2006, **128**, 13901; R. J. Trovitch, E. Lobkovsky, E. Bill and P. J. Chirik, *Organometallics*, 2008, **27**, 1470.
- [12] P.-H. Phua, L. LeFort, J. A. F. Boogers, M. Tristany and J. G. de Vries, *Chem. Commun.*, 2009, 3747; C. Rangheard, C. de J. Fernández, P.-H. Phua, J. Hoorn, L. Lefort and J. G. de Vries, *Dalton Trans.*, 2010, **39**, 8464.
- [13] K. Kano, M. Takeuchi, S. Hashimoto and Z. I. Yoshida, *J. Chem. Soc., Chem. Commun.*, 1991, 1728; M. Takeuchi and K. Kano, *Organometallics* 1993, **12**, 2059.
- [14] E. C. Ashby and J. J. Lin, *J. Org. Chem.*, 1978, **43**, 2567.
- [15] T. S. Carter, L. Guet and S. P. Thomas, *Chem. Commun.*, submitted.
- [16] A. Fürstner, R. Martin, H. Krause, G. Seidel, R. Goddarda and C. W. Lehmann, *J. Am. Chem. Soc.*, 2008, **130**, 8773; B. Bogdanović and M. Schwickardi, *Angew. Chem. Int. Ed.*, 2000, **39**, 4610.
- [17] A. G. Myers, B. H. Yang and D. J. Kopecky, *Tetrahedron Lett.*, 1996, **37**, 3623.
- [18] S. L. Buchwald and C. Bolm, *Angew. Chem. Int. Ed.*, 2009, **48**, 5586.
- [19] See supporting information for details.
- [20] J. A. Bull, J. J. Mousseau and A. B. Charette, *Org. Lett.*, 2008, **10**, 5485.
- [21] W. M. Czaplik, S. Grupe, M. Mayer and A. Jacobi von Wangelin, *Chem. Commun.*, 2010, **46**, 6350.